EXERGY AND ENTROPY ANALYSIS OF TRANSCRITICAL CO$_2$ REFRIGERATION SYSTEM

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Abstract. *The second law of thermodynamics deals with the quality of energy. More specifically, it is concerned with the degradation of energy during a process, the entropy generation, and the lost opportunities to do work; and it offers plenty of room for improvement. The aim of this paper is identify key factors that affect system performance with CO$_2$. Due to the impact of global warming of CFC´s and HFC's, the use of natural refrigerants has received worldwide attention. The natural refrigerant, carbon dioxide (CO$_2$) is promising for use in cooling systems, especially in the transcritical cycle. An exergetic analysis through the cycle of a bottle cooler adapted for commercial use with carbon dioxide was carried out so that the effectiveness of the system components can be estimated and classified, allowing direct efforts to improve performance of components of the transcritical cycle. The analysis revealed that the compressor and expansion valve are the largest sources of losses in the system, and therefore, efforts should focus on improving these components.*

Keywords: *exergy, entropy, carbon dioxide, transcritical cycle, natural refrigerants*

1. INTRODUCTION

The classical substances as hydrochlorofluorocarbons (HCFCs) and clorofluorcarbons (CFCs) used as working fluids in the vapour compression systems have to be replaced by new substances because of their ozone depletion potential and their greenhouse effect. Carbon dioxide (CO$_2$) is non-toxic, non-flammable, has zero ozone depletion potential and negligible global warming potential as refrigerant, excellent heat transfer coefficients, compatibility with material of refrigeration system and very low cost. At high ambient temperature, the carbon dioxide refrigeration cycle is operating in transcritical mode and at high working pressure because of the specific thermodynamic properties of carbon dioxide (relatively low critical temperature and relatively high critical pressure), (Yoon *et al.*, 2004).

This paper developed three scientific concepts: exergy, environment and sustainable development. The relations between exergy and both energy and the environment makes ir clear that exergy is directly related to sustainable development (Dincer and Rosen, 2007). Many suggest that mitigating the environmental impact of energy resource utilization and achieving increased resource utilization efficiency are best addressed by considering exergy. By extension, since these topics are critical elements in achieving sustainable development, exergy also appears to provide the basis for developing comprehensive methodologies for sustainability. The exergy of an energy form or a substance is a measure of its usefulness or quality or potential to cause change.

Industrial ecology is an approach to designing industrial systems that promotes systems that are less damaging to the environment. Jelinsky *et al.* (1992) defines industrial ecology as a new approach for the design of products and processes and to implement strategies for sustainable industrial production systems. It is a concept that industrial systems are seen in interaction with the environment that surrounds it. The objective is to optimize the cycle of materials. Within this concept there is the focus turned to the "Design for Environment" within which they can draw a restructuring environment industrial systems, and search for sustainable production systems. Industrial ecology methods can beneficially incorporate exergy to provide more powerful tools. Exergy methods are important since they are useful for improving efficiency.

According to Yumrutas *et al.* (2002) thermodynamic processes in refrigeration systems release large amounts of heat to the environment. Heat transfer between the system and the surrounding environment takes place at a finite temperature difference, which is a major source of irreversibility for the cycle. The first law of thermodynamics deals with the quantity of energy and asserts that the energy cannot be created or destroyed. This law merely serves as a necessary tool for the quantification of energy during a process and offers no engineering challenges. The second law, however, deals with the quality of energy. Çengel and Boles (2005) argue that this law has proved to be a very powerful tool in the optimization of complex thermodynamic systems. More specifically, she is concerned with the degradation of energy during a process, the entropy generation, and the lost opportunities to do work. Exergy based second law analysis of various systems has become a very effective tool to measure system effectiveness and to design the system to maximize energy savings. Exergetic analysis of a complex system can be performed by analyzing the components of the system separately. Identifying the main sites of exergy destruction, shows the direction for potential improvements. As you know all real thermodynamic processes are irreversible. Recently, Yumrutas *et al.* (2002) developed a methodology for exergetic analysis of vapour compression refrigeration systems, indicating that the major causative

factors of irreversibility in the refrigeration cycle include: friction and heat transfer across a finite temperature difference in the evaporator, compressor, condenser, refrigerant lines, pressure drops, subcooling and superheating, and heat gains in refrigerant lines. Brown and Domanski (2003) has proposed entropy generation methodology to compare the amounts of entropy generation in each of the primary system components, the analyses show that the CO_2 residential air-conditioning system generate 63% more entropy than the R-22 residential air-conditioning system. The biggest contributors to the lower CO_2 performance are irreversibilities associated with the heat rejection process and the expansion related processes. Aprea and Maiorino (2009) has proposed experiments comparing a commercial available R134a refrigeration plant to a cold store and a prototype R744 system working as a classical "split systems" to cool air in residential applications. The thermodynamic analysis shows that the component with the smallest exergetic efficiency is the expansion valve. In the transcritical cycle, the expansion valve contributes to the most significant part of the total exergy loss in the system, approximately 40%. Besides the exergetic analysis in the transcritical cycle shows an exergy loss in each component greater than the exergy loss of the corresponding components in a sub-critical refrigeration cycle.

Based on the above experimental and simulation work, the aim of this work is to identify the main factors that affect CO_2 system performance theoretically. For the analysis in question, will be used to cycle cooling system proposed by Cecchinato *et al.* (2006), which deals with the development of an R744 commercial single door bottle cooler that is cost competitive and matches the performance of a typical cost optimised R134a system. This paper is an case study of the second law applied to experimental results.

2. THE BOTTLE COOLER TRANSCRITICAL CYCLE

The transcritical refrigeration system proposed by Cecchinato *et al*. (2006) and object of analysis of this paper is a bottle cooler of 466 litres internal volume and 556 external volume. Figure 1 shows the tested bottle cooler.

This refrigeration system was re-designed to be adapted to CO_2 refrigerant. In this first realisation the main purpose was to verify the possibility of using components easily available on the market, yet assuring the same energy efficiency as the traditional appliances. The data of operation of the cycle were obtained through the utilisation of the test procedure for continuous running test for household applications, where the compressor runs continuously, at least for 24 hours, until thermal equilibrium is established. Tests were also carried out with an internal electric load of 188 W to keep the compartment temperature of CO_2 cabinet at 5°C. The ambient temperature was fixed in 30°C and 55% relative humidity. Table 1 shows the results from the transcritical cycle proposed by Cecchinato *et al*. (2006).

Table 1. Values of the main variables of transcritical cycle.

Variables	Unit	CO_2
Heat load from electrical power	W	188
Room temperature	^{o}C	28,8
Compartment mean temperature	^{o}C	5,2
Suction pressure	bar	32,0
Evaporation sat. temperature (at compressor suction)	^{o}C	-3,2
Discharge pressure	bar	91,7
Condensation sat. temperature (at compressor discharge)	^{o}C	-
Discharge temperature	^{o}C	93,4
Gas cooler outlet temperature	^{o}C	33,7
Internal HEX high pressure outlet temperature	^{o}C	25,1
Evaporator outlet temperature	^{o}C	4,1
Internal HEX vapour outlet temperature	^{o}C	18,0
Suction temperature	^{o}C	18,8
Approach at gas cooler	^{o}C	4,9
Superheating at evaporator	^{o}C	7,3
Mass-flow rate	kg/h	8,7
Cooling capacity	W	417
Compressor input energy	W	293
Compression isentropic efficiency	-	0,43
COP	-	1,422

Figure 1. The tested bottle cooler

Complementary computational tool for analysis of the system, the software CoolPack, is used for obtaining the P-h diagram and thermodynamic properties. In order to verify the results obtained by Cecchinato *et al.* (2006), the simulation of this cycle was performed with this software. Figure 2 shows the configuration of the transcritical cycle. The diagram of the cycle was obtained by the software, entering the parameters of the proposed system.
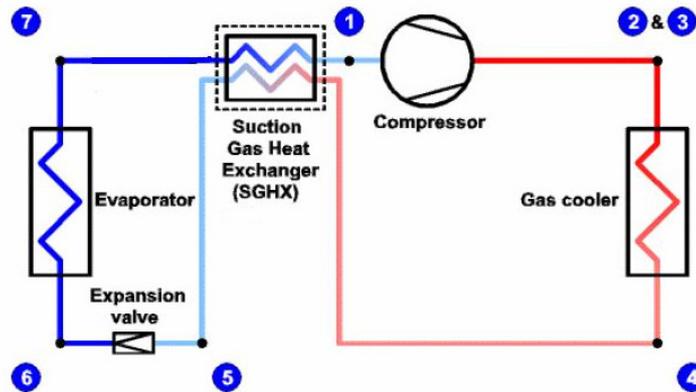
Figure 2. The bottle cooler transcritical cycle

From state point (1) the suction gas is compressed by the compressor and discharged (2) & (3). From the compressor the refrigerant is lead through a gas cooler. The refrigerant in the gas cooler outlet (4) is then lead to the "hot side" inlet of the suction gas heat exchanger (suction gas side is "cold side"), if this has been selected, or directly to the inlet of the expansion valve. From the "hot side" outlet of the suction gas heat exchanger (5) the refrigerant is expanded (6) into the evaporator. The evaporated and superheated refrigerant in the evaporator outlet (7) is then lead through the suction line to the "cold side" inlet of the suction of compressor (1). Figure 3 shows the P-h diagram of the bottle cooler cycle.
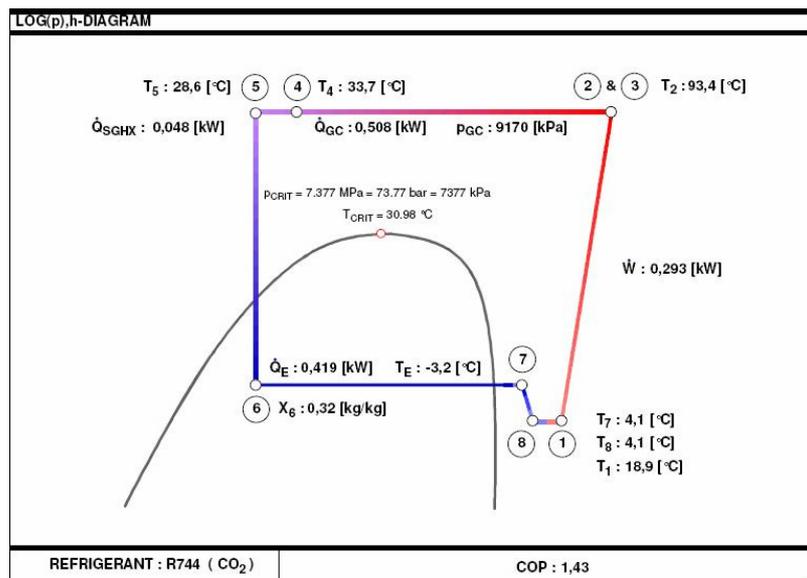

Figure 3. P-h diagram

In the Figure 3, was not considered the existence of the pressure loss between state points 7 and 1, and the existence of state point 8 was neglected, it is another heat exchanger that has been neglected in this analysis. The simulation using the CoolPack revealed that the compressor requires 293 W of work and 204 W lost to the environment in the form of heat (heat loss). Thus, the process of compression requires 89 W liquid power, this value has been used to determine the exergy loss of the compressor. Table 2 shows simulation results in CoolPack.

Table 2. Simulation results with the software.

Variables	Cecchinato *et al.*(2006)	CoolPack	Error (%)
Cooling capacity	417 W	419 W	0,50
Heat transfer in the gas cooler	-	508 W	-
Compressor heat loss	-	204 W	-
Heat transfer in the SGHX	-	48 W	-
COP	1,422	1,430	0,56

The operating parameters obtained by use of software have little disagreement with respect to Cecchinato *et al.* (2006), however the error is negligible, which validates the application of computational tool in the analysis of thermal systems for cooling. Table 3 shows the thermodynamic properties of CO_2 from CoolPack. The properties enthalpy and entropy calculated with the software will not be equal to those obtained with other programs, or taken from property any plots, if these programs and plots use a different reference condition.

Table 3. Thermodynamic properties of state points of transcritical cycle.

State point	T (^{o}C)	m (kg/s)	x (%)	P (kPa)	h (kJ/kg)	s (kJ/kgK)
1	18,9	0,00241	-	3199	463,69	1,93139
2/3	93,4	0,00241	-	9170	500,49	1,97322
4	33,7	0,00241	-	9170	290,08	1,28851
5	28,6	0,00241	-	9170	270,16	1,20487
6	-3,2	0,00241	32,0	3199	270,16	1,24987
7	4,1	0,00241	-	3199	443,89	1,90426

3. THERMODYNAMIC ANALYSIS

For performing a thermodynamic analysis for CO_2 cycle, the first law of thermodynamics is applied to the system. The first law of thermodynamics is a statement of the law of conservation of energy that readily yields the energy balance of the system,

$$Q_e + W = Q_g \tag{1}$$

were subscripts e and g represent the evaporator and gas cooler, respectively. The energy analysis of the cycle leads to computation of the Coefficient of Performance (COP), which indicates how efficiently energy is being used in the system. The COP is defined as the ratio of the desired heat transfer at the evaporator to the required work of the compressor,

$$COP = \frac{Q_e}{W} \tag{2}$$

From the conservation of energy, the work and heat transfer at each component in a stationary open system can be determined by

$$Q - W = \Delta H \tag{3}$$

The exergetic analysis will be performed based on the state points in table 3. These state points are distinguished by of the determination of properties such as pressure, temperature and/or quality. The system was modeled based on exergy balance (second law) of the individual components of the system. To simplify the analysis were taken three hypothesys: compression process is adiabatic but not isentropic, isoenthalpic expansion, pressure drops in all the connecting pipes and heat transfer between the connecting pipes and the ambient are negligible.

The main purpose for an exergy analysis is to determine where and how losses occur so that energy may be used most effectively. The definition of exergy a pure substance is given by

$$\psi = (h - h_0) - T_0(s - s_0) \tag{4}$$

The terms h_0 and s_0 represent values of enthalpy and entropy at the environmental temperature T_0. Exergy is a combined property of the system and its environment. Dincer and Rosen (2007) argue that the properties of the reference environment, determine the exergy of a system or stream. Thus, it is of utmost importance define the condition of reference to simplify the mathematical model. For the present calculation it is taken as having a temperature of 298.15 K and a pressure of 101.3 kPa. The choice of reference temperature of 298.15 K simplifies the mathematical model, because the enthalpy and entropy of CO_2 are zero in this condition (Srinivasan *et al.*,2003).

For each component will be determined exergy loss, defined as

$$\Delta \psi = \text{exergy supplied} - \text{exergy recovered} \tag{5}$$

and its exergetic effectiveness by

$$\xi = \text{exergy recovered} / \text{exergy supplied} \tag{6}$$

Based on above, below follows the mathematical modelling in all components of cycle, in accordance with methodology suggested by Fartaj *et al.* (2004).

3.1. Compressor

The exergy supplied in the compressor is given by

$$E_{in} = \overset{\bullet}{W} = \overset{\bullet}{m} \cdot (h_2 - h_1) \tag{7}$$

The exergy recovered in the compressor is given by

$$\psi_C = \overset{\bullet}{m} \cdot (\psi_1 - \psi_2) \tag{8}$$

The exergy loss in the compressor is given by

$$\Delta \psi_C = \psi_1 + \overset{\bullet}{W} - \psi_2 = E_{in} - \psi_C = \overset{\bullet}{m} \cdot (T_0 (s_2 - s_1)) \tag{9}$$

Where $\overset{\bullet}{m}$ is the mass flow rate of refrigerant in the cycle.

3.2. Evaporator

The exergy supplied in the evaporator is given by

$$\psi_S = \overset{\bullet}{Q}_E \left(1 - \frac{T_0}{T_r} \right) \tag{10}$$

The exergy recovered in the evaporator is given by

$$\psi_R = \overset{\bullet}{m} \cdot (\psi_7 - \psi_6) \tag{11}$$

The exergy loss in the evaporator is given by

$$\Delta \psi_{ev} = \psi_7 + \overset{\bullet}{Q}_E \left(1 - \frac{T_0}{T_r} \right) - \psi_6 = \overset{\bullet}{m} \cdot (h_7 - T_0 s_7) + \overset{\bullet}{Q}_E \left(1 - \frac{T_0}{T_r} \right) - \overset{\bullet}{m} \cdot (h_6 - T_0 s_6) \tag{12}$$

The term $\overset{\bullet}{Q}_E \left(1 - \frac{T_0}{T_r} \right)$ is thermal exergy loss rate, where $\overset{\bullet}{Q}_E$ is heat absorbed in evaporator and T_r is temperature at the boundary of component. According with Arora and Kaushik (2008) the thermal exergy loss rate is related to external irreversibility which takes place because of temperature difference between the control volume and the surroundings. If the system includes the immediate surroundings then the boundary of the thermal system is at the same temperature as the temperature of immediate surroundings and hence the value of thermal exergy loss turns out to be zero.

3.3. Expansion valve

$$\Delta \psi_V = \overset{\bullet}{m} \cdot (\psi_6 - \psi_5) = \overset{\bullet}{m} \cdot (T_0 (s_5 - s_6)) \tag{13}$$

3.4. Gas cooler

The exergy supplied in the gas cooler is given by

$$\psi_S = \overset{\bullet}{m} \cdot (\psi_2 - \psi_4) \tag{14}$$

The exergy recovered in the gas cooler is given by

$$\psi_R = \overset{\bullet}{Q} \left(1 - \frac{T_0}{T_{GC}} \right) \tag{15}$$

The exergy loss in the gas cooler is given by

$$\Delta \psi_{GC} = \psi_2 - \overset{\bullet}{Q}_g \left(1 - \frac{T_0}{T_{gc}} \right) - \psi_4 = \overset{\bullet}{m} \cdot (h_2 - T_0 s_2) - \overset{\bullet}{Q}_g \left(1 - \frac{T_0}{T_{gc}} \right) - \overset{\bullet}{m} \cdot (h_4 - T_0 s_4) \tag{16}$$

The gas cooler heat rejects to ambient by an amount $\overset{\bullet}{Q}_g$ at variable temperature (glide). For heat transfer occurring at variable temperature, Fartaj *et al.* (2004) uses the entropic average temperature. The entropic average temperature is often used to streamline and simplify the analyses of processes involving variable temperature heat transfer. It is defined as

$$T_{gc} = \frac{h_2 - h_4}{s_2 - s_4} \tag{17}$$

It is assumed that the heat $\overset{\bullet}{Q}_g$ is rejected to the cooling air at the constant temperature T_{gc}.

3.5. Internal heat exchanger (SGHX)

The exergy supplied in the internal heat exchanger is given by

$$\Delta \psi_S = \overset{\bullet}{m} \cdot (\psi_4 - \psi_5) \tag{18}$$

The exergy recovered in the internal heat exchanger is given by

$$\Delta \psi_R = \overset{\bullet}{m} \cdot (\psi_1 - \psi_7) \tag{19}$$

The exergy loss in the internal heat exchanger is given by

$$\Delta \psi_{HEX} = \overset{\bullet}{m} \cdot ((\psi_4 - \psi_5) - (\psi_1 - \psi_7)) = \overset{\bullet}{m} \cdot ((h_4 - h_5 + T_0 (s_5 - s_4)) - ((h_1 - h_7) + T_0 (s_7 - s_1))) \tag{20}$$

To complement the analysis of irreversibilities in the transcritical cycle, is also an analysis performed by the entropy balance method in all components of the cycle. For the transcritical cycle operating with CO_2, the entropy balance method can be deduced as follows

$$\sum \delta S_i = \frac{Q_{gc}}{T_{agc}} - \frac{Q_e}{T_{ae}} \tag{21}$$

where δS_i is the irreversible entropy increase in the system, T_{agc} and T_{ae} are the average temperatures of the surroundings of gas cooler and evaporator, respectively, Q_{gc} and Q_e are the amounts of heat transfer in the gas cooler and evaporator. For each component, the irreversibilities can be identified

$$\text{Gas cooler, } \delta S_{gc} = \frac{Q_{gc}}{T_{agc}} - \frac{Q_{gc}}{T_{gc}} \tag{22}$$

$$\text{Evaporator, } \delta S_e = \frac{Q_e}{T_e} - \frac{Q_e}{T_{ae}} \tag{23}$$

$$\text{Compressor, } \delta S_c = s_2 - s_1 \tag{24}$$

$$\text{Expansion valve, } \delta S_{ev} = s_6 - s_5 \tag{25}$$

$$\text{Internal heat exchanger, } \delta S_{sghx} = (s_5 + s_1) - (s_4 + s_7) \tag{26}$$

In the Eqs. 22 at 27, the first term of the difference represents the supplied entropy and the second term represents the recovered entropy. The values of the energy (kJ/kg) in the gas cooler and evaporator in the Eqs. 22, 23 and 24 represents the values of the table 4. The terms T_e and T_{gc} represents, respectively, the evaporation temperature and the entropic average temperature in the gas cooler.

4. RESULTS AND DISCUSSIONS

Table 4 shows the analysis of the transcritical cycle with the first law. Energy and mass balances were written around each components and combined with the thermodynamic properties to yield the results listed in the table using the Eq. (1), Eq. (2) and Eq. (3). The sum of the refrigeration capacity (heat input in the evaporator) and the work consumed by the compressor is equal to the heat rejected at the gas cooler. The COP is defined as the ratio between cooling capacity and the work consumed in the compressor. In this case, COP = 1,43. It is noteworthy that for calculation of the COP was considered the work as being consumed in the compressor of 293 W, which represents the total energy provided in this component, 204 W wasted in the form of heat loss and 89 W for the refrigerant compression.

Table 4. First law analysis of the transcritical cycle

Component	Energy (kJ/kg)	Energy Flow (W)	COP
Compressor	36,80	88,688	-
Gas Cooler	210,67	507,714	-
Evaporator	173,730	418,689	-
Internal heat exchanger	19,660	47,381	-
System	-	-	1,430

The results of exergetic analysis are presented in Table 5. Column 1 shows the components of the order of the largest exergy loss (column 4). Columns 2 and 3 list the values of exergy supplied and recovered for each component of the transcritical cycle. An important parameter of evaluation is the exergetic effectiveness, which is defined as the ratio between the exergy recovered and exergy supplied. This parameter shows the difference between the performance of each component with respect to the ideal case. This result allows attention to be concentrated on the aspects of system operation that offer the greatest opportunities for improvement of the system performance (Fartaj *et al*, 2004).

Analyzing the results, it appears that the compressor and expansion valve have the highest rates of exergy loss, approximately 60% of the total exergy loss. To improve performance of the cycle, special attention should be directed to the improvement these components, reducing the irreversibilities associated with the processes of compression and expansion, as the irreversibility lead to greater exergy destruction. The exergetic effectiveness indicates the degree of difficulty of improvement in the components. The compressor has more capacity for improvement in relation to the expansion valve, since the efficiency of the compressor is 66,0% and the valve is 86,9%.

Because of the high pressure drop occurring in the system, the expansion device contributes a much larger fraction of the irreversibility compared to conventional systems (subcritical cycle). In the compressor, process irreversibility, pressure loss due to friction in inlet and outlet valves and heat loss to the environment are the basic reasons for the exergy loss in this component (Sarkar et al., 2005). In this evaluation, the heat loss through of the compressor, representing approximately 70% of total work supplied, is the main cause of high rate of exergy destruction in this component, being influenced also the low isentropic efficiency of the compressor Danfoss TN1406. From viewpoint of the energetic efficiency, Cecchinato et al. (2006) says that the CO_2 shows the worst performance when compared with the other two synthetic fluids: R134a and R404a. This worsening of energy efficiency is related to exergy losses during cyclic running, that penalise this fluid more than the competitors.

Looking at the gas cooler and evaporator, the irreversibility in components lead to substantial rates of exergy loss, mainly in the evaporator, occur mainly due to the difference in temperature between the heat exchangers and the surroundings. The effective temperature of the fluid in heat exchangers can be reduced and reduce consequently the irreversibility in these components, by increasing the area of heat transfer, fins may be incorporated or increase in the total length of the heat exchanger. The bottle cooler was re-designed to be adapted to R744 refrigerant, the original condensing finned coil was replaced by a more efficient one acting as gas cooler, having 2,69 m^2 external heat transfer area. The increase in area, approximately +32%, was obtained with only +11% increase in depth dimension and with the same face area. Two main reasons are considered to explain the low exergy loss of the gas cooler, the first relates to the entropic average temperature in this component is located close to environmental temperature, reducing the rate of irreversibility, the second is the improvement of the gas cooler used the experiment, approaching the optimum performance for this type of equipment. The internal heat exchanger/suction gas heat exchanger (SGHX) presented considerable exergy loss, this fact can be attributed to the difference in temperature between the sides of high and low pressure. Other relevant information is related to a constructive form of heat exchanger, because as cooling system is relatively small, the heat exchanger was obtained through physical contact between the pipes for high and low pressure with insulating material, it is important analyze the efficiency of this constructive form to achieve greater effectiveness of this component.

Table 5. Analysis of the transcritical cycle by exergy method

Component	Exergy supplied (W)	Exergy recovered (W)	Exergy loss (W)	Effectiveness (%)
Compressor	88,688	58,618	30,07	66,0
Expansion valve	246,996	214,663	32,33	86,9
Gas cooler	13,700	12,900	0,870	94,1
Evaporator	51,517	44,020	7,497	85,4
Internal HEX	28,223	12,091	16,132	42,8
System	-	-	86,899	-

Figure 4 shows the exergy flow diagram, the consumption of exergy in the components of the transcritical cycle, illustrates the degradation of available energy (exergy) through the system. Can be observed that the compressor and expansion valve have the greatest losses of the total exergy loss of the system. The percentage values in the diagram were obtained through the relationship between exergy loss at each component and exergy supplied in the system (88,688 W). In a real process there will always be destructions and losses resulting in a greater exergy supplied into the process on the exergy of the product.
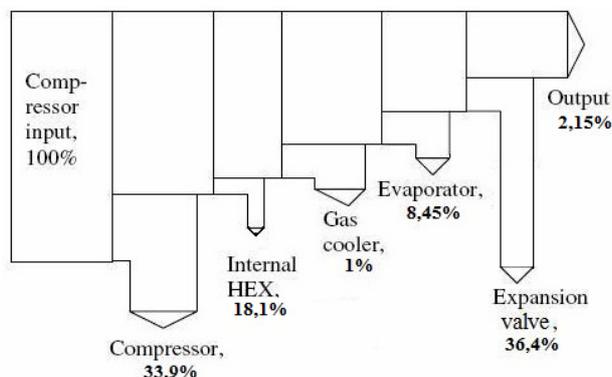

Figure 4. Exergy flow diagram of the bottle cooler transcritical cycle

Table 6 and Figure 5 presents the results obtained from entropy method. The results were obtained with Eq. (22) at the Eq. (27). The fourth column lists the entropy increase in each component due to irreversibilities that exist in the real processes. The last column lists the contributions that are defined as the component entropy gain as a percentage of the total entropy gain in the system.

Table 6. Analysis of the transcritical cycle by entropy method

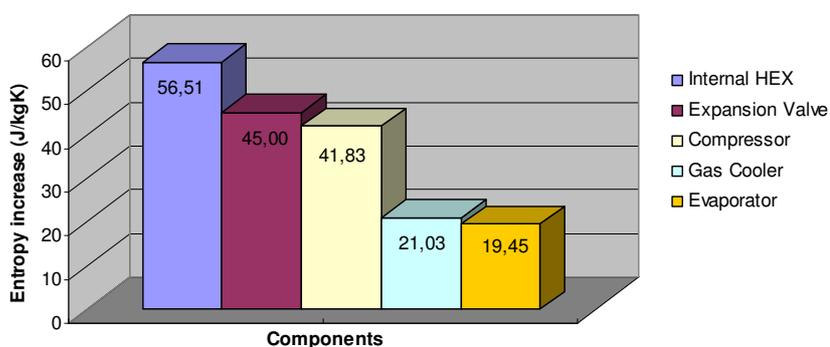| Component	Entropy supplied (kJ/kgK)	Entropy recovered (kJ/kgK)	Entropy increase (kJ/kgK)	Contribution (%)
Compressor	1,97322	1,93139	0,04183	22,7
Expansion valve	1,24987	1,20487	0,04500	24,4
Gas cooler	0,70659	0,68555	0,02103	11,4
Evaporator	0,64392	0,62447	0,01945	10,5
Internal HEX	3,19277	3,13626	0,05651	30,7
System	0,70659	0,62447	0,08212	100,0

Figure 5. Entropy increases on the system components

The results show that there is a discrepancy between the entropy increase for the system with the sum of the entropy in each components. This fact can be attributed to entropy generation that occurs in components (pressure drop, for example) that is not taken into account in the calculation through the system. The results for each component have the same trends of results for the method of exergy loss, where the expansion valve, compressor and heat exchanger internal present the highest values of entropy increase. Since the compressor and expansion valve are the bottlenecks for the development of this experimental system for improvements in the overall efficiency. Yang *et al.* (2005) performed an comparative study for the transcritical carbon dioxide refrigeration cycles with a throttling valve (without internal HEX) and with expander. It is found that ih the throttling valve cycle, the largest exergy loss occurs in the throttling valve, about 38% of the total cycle irreversibility With the application of the expander, the irreversibility loss of the expansion process is reduced almost by 50%. The authors considered the isoentropic efficiency of the compressor as 75%, to 43% in this paper. Robinson and Groll (1998) presented two thermodynamic models for transcritical carbon dioxde cycles with and without an expansion turbine. The efficiencies of the transcritical carbon dioxide cycles will be compared with the ones of conventional R-22 cycles. The expansion valve is the component with the largest percentage of total irreversibility in the expansion valve carbon dioxide cycle. Replacing the expansion valve with an expansion work recovery turbine with an isentropic efficiency of 60% reduces the process´s contribution to total cycle irreversibility by 35%. Brown *et al.* (2002) evaluated performance merits of CO_2 and R134a automotive air conditioning systems. The entropy generation calculations indicated that the large entropy generation in the gas cooler was the primary cause for the lower performance of CO_2. The large CO_2 temperature glide (approximately 80 versus 25^oC for R134a) and the larger amount of heat needed to be rejected by the CO_2 gas cooler than by the R134a condenser causes a significant amount of entropy generation.

5. CONCLUSIONS

The results of the simulation and second law analysis of transcritical bottle cooler system operating with CO_2 led to the following conclusions :

a) The exergy analysis applied in this work is not used from the viewpoint of evaluation of the environmental impacts, because it is a system of small size and unitary. This method was applied to detect and classify the irreversibilities and exergy losses that occur in the components of the transcritical cycle in order to promote improvements.

b) The mathematical model that was presented shown to be compatible with the system based on empirical data proposed by Cecchinato *et al.* (2006). The results show that directions should be taken to improve the experimental model.

c) The complementary simulation results with the software CoolPack show that 204 W is wasted as heat by the compressor, approximately 70% of the work consumed by this component.

d) Compressor, expansion valve, internal heat exchanger and evaporator contribute to system exergy loss to a larger extent, while the gas cooler has a negligible effect. The expansion valve presents greater exergy loss because of the high pressure drop occurring in the system. Capillary tubes can be a good engineering option for small capacity systems in lieu of an expansion valve, which has been thought of as the only possible solution to attain the pressure optimization, an important feature of all transcritical CO_2 systems. The system exergetic efficiency is better with a capillary tube than that of a controllable throttle valve system (Agrawal and Bhattacharyya, 2009). The internal heat exchanger/suction gas heat exchanger (SGHX) presented considerable exergy loss, this fact can be attributed to the difference in temperature between the sides of high and low pressure.

e) Compressor, expansion valve and internal heat exchanger have the largest contributions to the entropy increase of the system. The compressor presents the largest capacity of improvements with respect to the expansion valve, since as presented exergetic effectiveness of 66%, being more susceptible to improvements than the expansion valve.

f) The review of literature consolidates the compressor and expansion valve as the main sources of exergy losses in the transcritical cycle. Some authors indicates the gas cooler as larger irreversibility focus, such not verified in this paper. Was demonstrated the key factors on the research in the commercial light refrigeration systems for CO_2.

g) Efforts should focus primarily on improving the compressor and expansion valve. To improve the COP of the transcritical cycle of the R744 it is necessary both to optimize the components and to use some arrangements as the recovery of the compression work and internal heat exchanger.

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